

Preparation, Characterization of A New Cis – Iodoplatin and Cis - Carboplatin Complexes and Study There Spectral, Physical and Pharmaceutical Properties

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Abstract

This search includes preparation of new cis-platin, in the form of iodine, and new two cis - carboplatin complexes {cis [Pt(4-aminoantipyrine) (NH₃)₂]I₂ and cis[Pt (aminosalicylic acid) (NH₃)₂]I}, which used as anti - cancer drugs, they are prepared by using the organic ligands 4-amino salicylic acid and 4-amino anti pyrine. These complexes were characterized by using UV-Vis., FTIR, XRD and ¹HNMR spectroscopic techniques . The effect of the concentration and the temperature on the molar electrical conductivity of its solutions were studied. The expiry date (half-life) for the solutions of the complexes in the circumstances at the degrees (35⁰C) and (20⁰C) was determined.

Keyword: Cis-[Diaminodiodoplatinum(II)], Carboplatin, XRD, ¹HNMR.

Abbreviations:

4 – aminoantipyrine: 4-AAP

4 – aminosalicylic acid: 4-ASA

Introduction

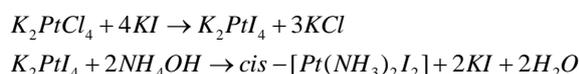
Cis – platin and cis - carboplatin complexes are coordination complexes of platin (II) and classify as anti - cancer drugs according to American drugs record[1]. The DNA nuclear acid in the cell is the target of the cis – platin [2],[3]. Cis - carboplatin is less toxic than cis - platin and more active against the same tumor[4]. Carboplatin had been using for long time for the treatment of many types of tumors such as lung tumors and contrast of cis-platin mechanism, carboplatin linked

covalently with the DNA nuclear acid and do not cause any kidney poisoning or nausea or puke and its poisoning is inhibition bone marrow, which leads to decrease the platelets, therefore, it is more used than cis – platin for the patients who suffer from toxic intolerances of cis – platin, also cis - carboplatin has more stability in blood plasma than cis - platin [5]. The proposed mechanism for cis – platin to stop the cell division is: cis – platin can enter inside the cell easily and associated with the two sandwiches of the DNA at the same by hydrogen bonds during nitrogen, sulfur and oxygen atoms cause to kill the cancer cell[6].

Excremental

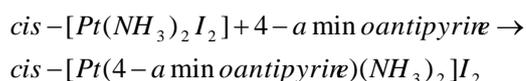
1-Preparation of cis – iodoplatin: [Pt(NH₃)₂I₂] [7-10]:

In a small beaker (0.003mol ,1.25g) of (K₂PtCl₄) was put and (2ml) of distilled water was added. The solution was heated in a water bath at (40 °C) and KI solution ((0.015mol)(2.5g) in 5ml hot distilled water was added. The color of the resulting solution changed from red-brown to red- black. This solution was heated to (70 °C) with stirring after that add drop wise (5ml) of ammonia solution (2M) to control the Ph at 7, soon a fine yellow crystals of cis- [Pt (NH₃)₂I₂] precipitate was formed. To adjust the PH, add ammonium chloride (0.2g) as a buffer solution to prevent the formation of hydroxide complex. Heat the solution to a (60 °C) for one hour and take care of that the pH not raised above 7. Cool the solution by using ice bath, soon greenish – brown precipitate was formed. Filter and wash with 10ml of cool ethanol and with 30 ml of cool distilled water, product percentage was (85%) according to the following equations



2-Preparation of cis - [Pt(4-aminoantipyrine) (NH₃)₂]I₂:

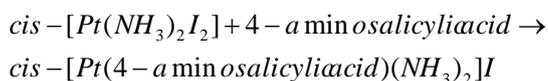
In distilled flask (0.00083mol,0.4g) of cis- [Pt (NH₃)₂I₂] and (0.00083mol,0.17 g) of (4-amino anti pyrine) reagent were mixed and dissolved in (20ml) of absolute ethanol. By reflux process, at (80 °C) with stirring for three hours, the reaction was done. Brown precipitate was formed. The mixture was cooled in ice bath. Filtrated and washed with (10ml) of ethanol and (20 ml) of cooled water. The following equation to make clear this reaction:



3-Preparation of cis - [Pt(4-aminosalicylic acid)(NH₃)₂]I₂:

In distilled flask (0.00083mol)(0.4g) of cis- [Pt (NH₃)₂I₂] and (0.00083mol)(0.122 g) of (4-amino salicylic acid) reagent were mixed and dissolved in (20ml) of

absolute ethanol. The pH of the solution was adjust between (8-7) so the solution become light basic by using very dilute potassium hydroxide solution. By reflux process, at (80 °C) with stirring for three hours, the reaction was done.. The mixture was cooled in ice bath. Filtrated and washed with (10ml) of ethanol and (20 ml) of cool water Black precipitate was formed. The following equation to make clear this reaction:



4 - Physical, spectral and pharmaceutical studies:

The prepared complex was characterized by UV. Vis., FTIR and HNMR spectra. XRD spectra for the cis – iodoplatinum complex was recorded by using X-ray diffraction technique. HNMR spectra for the complexes were measured. Molar conductivity was measured, for the prepared complexes with different concentrations, from 8×10^{-3} M to 1×10^{-4} M at 35 °C. The change in molar connectivity at different temperatures from 80 °C to 20 °C was measured. The effect of temperature of the prepared complexes (4×10^{-4} M) studied by using UV. Vis. technique at different temperatures (80-20 °C). The expiry date of the complex solutions (1×10^{-3} M) at two temperatures, 40 °C and 20°C, was studied by using UV. Vis. Technique, in which the absorbance was measured at each (48) hour for 20 days.

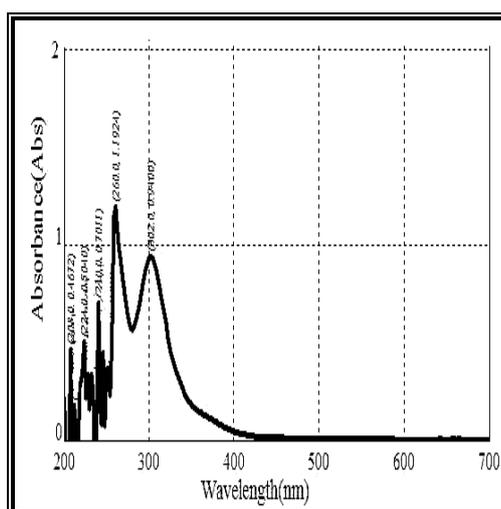
Results and Discussion

1-UV.Vis. spectra:

The electronic spectra of the complexes cis-[Pt(4-aminoantipyrine) (NH₃)₂] I₂ and cis-[Pt(4-aminosalicylicacid) (NH₃)₂]I, show that λ_{max} are (495 nm) (440 nm) respectively which belong to the transmission (M→L.C.T.) [11]. The λ_{max} (364nm) (330 nm) which belong to the transmission (n→π*) [12] and the λ_{max} (296 nm) (304 nm) which belong to the transmission (π→π*) [13], while the complex cis-[Pt(NH₃)₂I₂] shows a λ_{max} (302 nm) which belongs to the transmission (M→L.C.T.). table (1)and figures (1,2,3) show these results. The differences in λ_{max} between the primary material and the complexes is a good evidence for the formation the new complexes.

Table (1): UV.Visb. data for the prepared complexes

Compounds	λ (nm)	V(cm ¹)	Type of transition
Cis-[Pt(NH ₃) ₂ I ₂]	302	33112	M→L.C.T.
Cis-[Pt(4-AAP) (NH ₃) ₂] I ₂	296	33783	$\pi^* \rightarrow \pi$
	364	27472	n→ π^*
	495	20202	M→L.C.T.
Cis-[Pt(4-ASA) (NH ₃) ₂]I	304	23894	$\pi^* \rightarrow \pi$
	330	30303	n→ π^*
	440	22727	M→L.C.T.

Figure (1): UV-Vis. spectrum for cis-[Pt(NH₃)₂I₂]

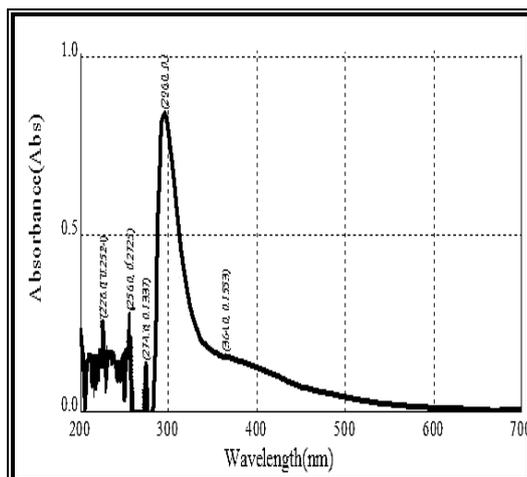


Figure (2): UV-Vis. spectrum for cis-[Pt(4-AAP) (NH₃)₂] I₂

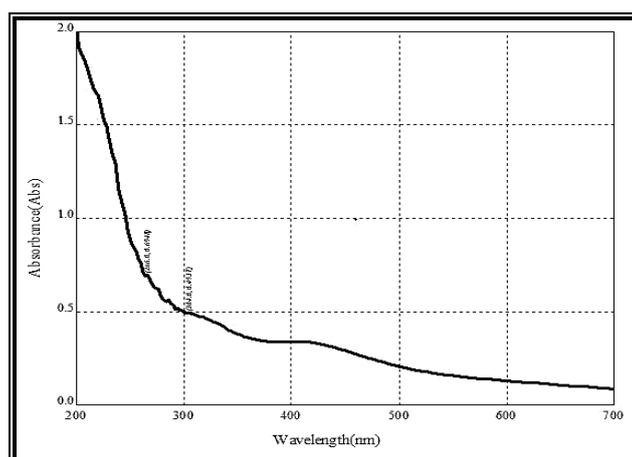


Figure (3): UV-Vis. spectrum for cis-[Pt(4-ASA) (NH₃)₂]I

FTIR spectra:

The FTIR spectra for the complexes cis-[Pt(NH₃)₂I₂], cis-[Pt(4-AAP)(NH₃)₂] I₂ and cis-[Pt(4-ASA) (NH₃)₂]I show the frequencies (595 cm⁻¹), (586 cm⁻¹) and (546 cm⁻¹) belong to the bond (M-N) respectively. The the frequencies (498 cm⁻¹) and (441 cm⁻¹) for the complexes Cis-[Pt(4-AAP)(NH₃)₂] I₂ and cis-[Pt(4-ASA) (NH₃)₂]I respectively belong to the transmission (M-O) respectively, and the frequency (3251 cm⁻¹) belongs to the bond (O-H) of the complex cis-[Pt(4-ASA) (NH₃)₂]I [14], table (2) and figures (4,5,6) show these results. The shifting,

removing and appearing of some frequencies is an evidence of the formation of new complexes.

Table (2): Explain the frequencies FTIR spectra for the complexes

Compounds	(N-H)	(N-N), (N=N)	(C=O)	(M-N)	(M-O)	(O-H)
Cis-[Pt(NH ₃) ₂ I ₂]	3272s 1290w			595 m		
Cis-[Pt(4-AAP)(NH ₃) ₂] I ₂	3431s 1161w	1492s	164 3s	586 m	498 m	
Cis-[Pt(4-ASA)(NH ₃) ₂]I	3461s 1028w		160 8s	547 m	441 m	325 1b

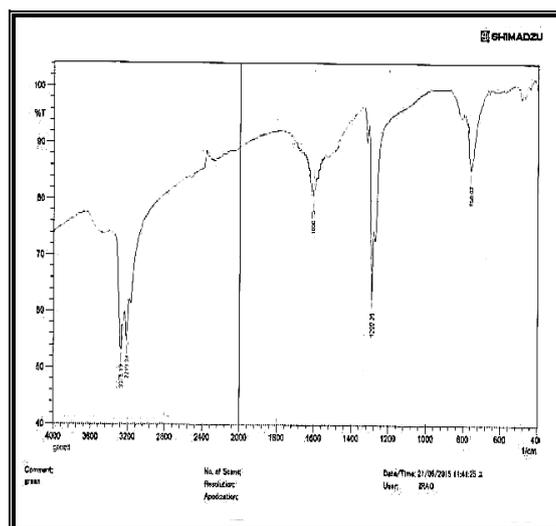


Figure (4): FTIR spectrum for Cis-[Pt(NH₃)₂I₂]

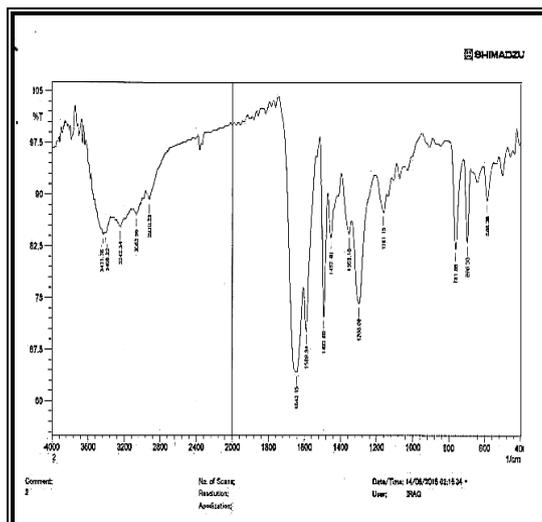


Figure (5): FTIR spectrum for Cis-[Pt(4-AAP) (NH₃)₂] I₂

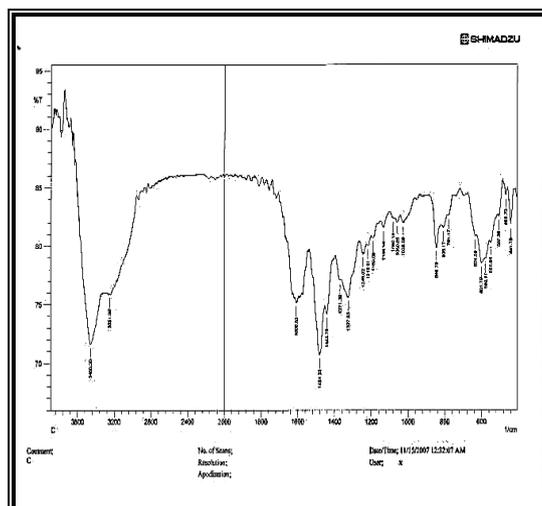


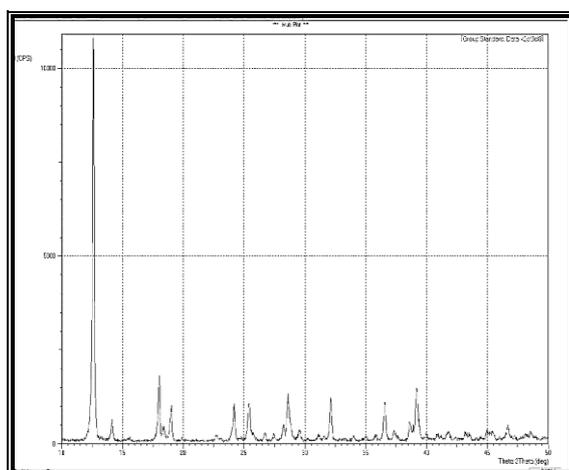
Figure (6): FTIR spectrum for Cis-[Pt(4-ASA) (NH₃)₂]I

XRD spectra:

X-ray diffraction angles(degree) for the complex cis-[Pt(NH₃)₂I₂] are 12.68, 18.06 and 39.19, the first value belongs to Pt(II), that ensures that the complex contains Pt(II) rather than Pt(IV) ion, which ensures what we want. Table (3) and figure (7) show these results.

Table (3): XRD data for the cis – iodoplatin complex

Compounds	2 Theta (deg.)	I/I ₂
(Cis-[Pt(NH ₃) ₂ I ₂])	12.68	100
	18.06	15
	39.19	14

**Figure (7): XRD spectrum for cis-[Pt(NH₃)₂I₂] complex**

¹HNMR Spectra

¹HNMR spectra of the complexes cis-[Pt(NH₃)₂I₂], cis-[Pt(4-AAP) (NH₃)₂] I₂ and cis-[Pt(4-ASA) (NH₃)₂]I had been studied. ¹HNMR spectrum of the complex cis-[Pt(NH₃)₂I₂] shows an appearance a single peak at 3.32 ppm which belongs to the six protons of the two groups of NH₃ as obvious in figure 8. ¹HNMR spectrum of the complex cis-[Pt(4-AAP) (NH₃)₂] I₂ shows an appearance of a single peak at 6.70 ppm which belongs to the benzene ring within the complex and two single peaks at 2.09 ppm and 2.70 ppm, the first belongs to CH₃ group that attached to carbon atom and the second belongs to CH₃ group that attached to nitrogen atom in pyrrol ring, additionally, there is a single peak at 3.8 ppm belongs to NH₂ group, figure 9 shows all these peaks. ¹HNMR spectrum of the complex cis-[Pt(4-ASA) (NH₃)₂]I shows an appearance of a single peak at 7.10 ppm which belongs to the aromatic ring, also there is a single peak at 3.33 ppm which belongs to NH₂ group, additionally, there

are two peaks at 2.74 ppm and 2.09 ppm which belong to NH₃ group. All these peaks proved that the ligands and the ion metal joined together to form the complexes.

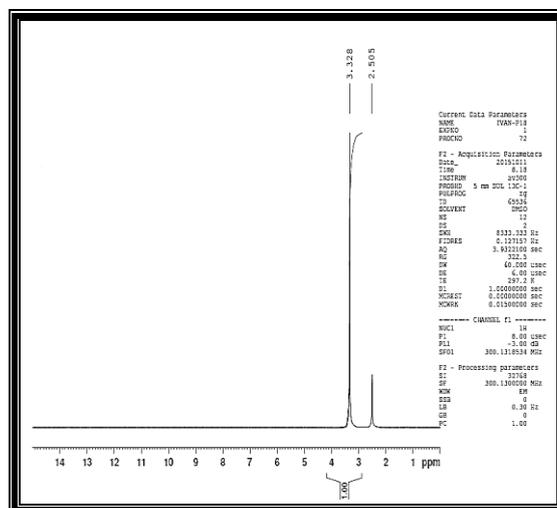


Figure (8): ¹H NMR spectra for cis-[Pt(NH₃)₂I₂] complex

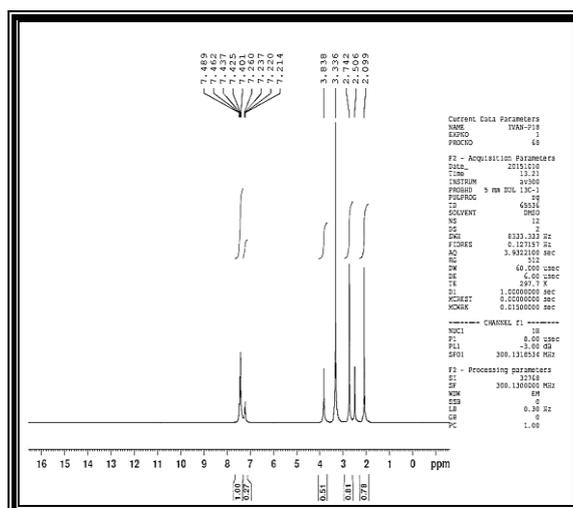


Figure (9): ¹H NMR spectra for cis-[Pt(4-AAP)(NH₃)₂] I₂ complex

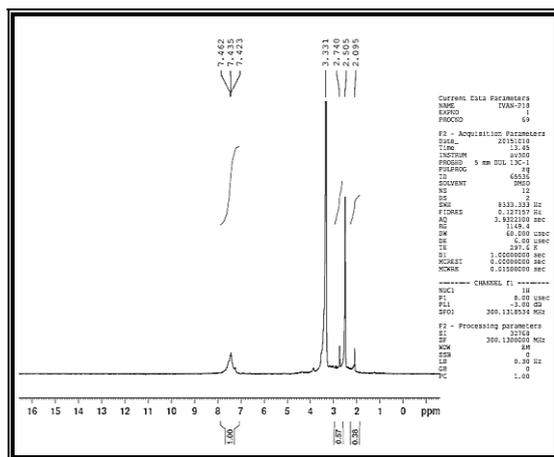


Figure (10): ^1H NMR spectrum for cis-[Pt(4-ASA) $(\text{NH}_3)_2$]I complex

The effect of concentration on molar conductivity:

The effect of concentration on the molar conductivity was studied for the solutions of the complexes. This study shows that the complexes cis-[Pt(4-AAP) $(\text{NH}_3)_2$] I₂ and cis-[Pt(4-ASA) $(\text{NH}_3)_2$]I have an increase in the molar conductivity with the increasing in the concentration of the complexes. Table (4) and figures (11,12) show the data of this study.

Table (4): the effect of conc. on molar conductivity for the complexes.

Con. (M)	Cond. Mc/cm Cis-[Pt(4-AAP) $(\text{NH}_3)_2$] I ₂	Cond. Mc/cm Cis-[Pt(4-ASA) $(\text{NH}_3)_2$]I
0.0001	2.24	38.8
0.0005	4.76	156.0
0.0010	8.45	278.0
0.0020	13.79	502.0
0.0030	18.82	722.0
0.0040	23.20	903.0
0.0050	27.60	1105.0
0.0060	32.00	1243.0
0.0070	36.30	1481.0
0.0080	39.50	1591.0

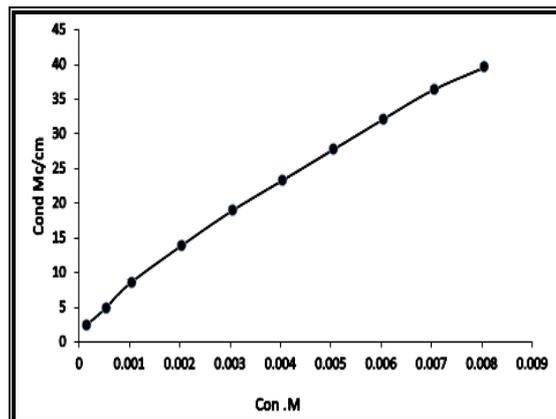


Figure (11): The effect of con. on the cond. for the complex $\text{cis-[Pt(4-AAP)(NH}_3\text{)}_2\text{]I}_2$

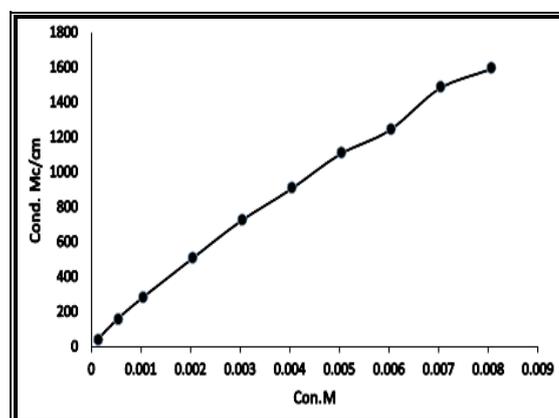


Figure (11): The effect of con. on the cond. for the complex $\text{cis-[Pt(4-ASA)(NH}_3\text{)}_2\text{]I}_2$

The effect of temperature on molar conductivity:

The effect of temperature on the molar conductivity of the complexes ($1 \times 10^{-3} \text{M}$) has been studied. The complexes $\text{cis-[Pt(4-AAP)(NH}_3\text{)}_2\text{]I}_2$ and $\text{cis-[Pt(4-ASA)(NH}_3\text{)}_2\text{]I}$ had shown that there is an increase in the molar conductivity with an increase the temperature. Table (5) and figures (13,14) show the results of this study.

Table (5): The effect of temp. on the molar conductivity of the complexes.

Temp. C°	Cond. Mc/cm Cis-[Pt(4- AAP)(NH ₃) ₂] I ₂	Cond. Mc/cm Cis-[Pt(4-ASA) (NH ₃) ₂]I
20	1.76	97.0
25	2.00	98.4
30	2.19	99.3
35	2.50	99.9
40	3.00	101.1
45	3.34	101.2
50	3.36	101.6
55	3.40	102.0
60	3.80	102.2
65	4.32	102.3
70	4.68	102.3
75	5.13	102.4
80	5.65	102.4

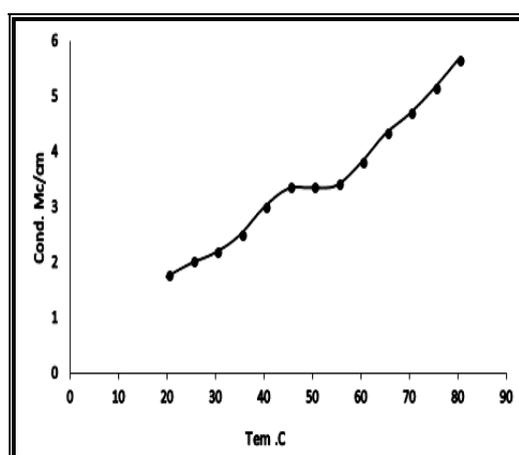


Figure (13): Effect of temp. on cond. cis-[Pt(4-AAP) (NH₃)₂] I₂

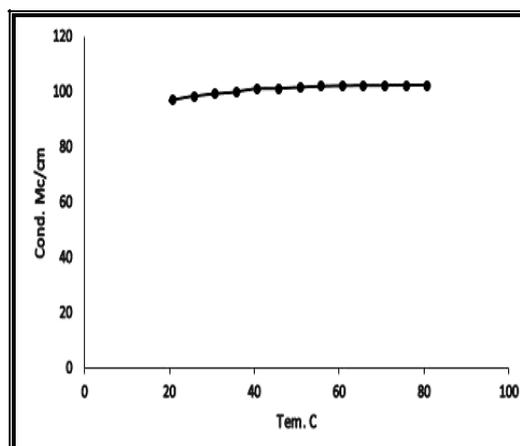


Figure (14): Effect of temp. on cond. cis [Pt(4-ASA) (NH₃)₂]I

The effect of temperature on the stability of the complexes:

The effect of temperature on the stability of the complexes had been studied by using UV.Vis. spectroscopy in which the absorbance of the solutions were measured. This study shows that the complex cis-[Pt(4 ASA) (NH₃)₂]I is stable until 50 °C and above this degree begins to break part, while the complex cis-[Pt(4AAP)(NH₃)₂] I₂ stable until 80 °C, that may be because of the composition structure of the ligands. Table (6) and figures (15, 16) show this study.

Table (6): the effect of temperature on the complexes stability

Tem. ^o C	Abs. Cis- [Pt(4AAP)(NH ₃) ₂] I ₂	Abs. Cis-[Pt(4 ASA) (NH ₃) ₂]I
20	1.6121	0.1320
25	1.6443	0.0999
30	1.6578	0.0961
35	1.6649	0.0960
40	1.6767	0.0538
45	1.6852	0.0445

50	1.6960	0.0405
55	1.7130	0.0237
60	1.7812	0.0166
65	1.8144	0.0136
70	1.9256	0.0030
75	1.9783	0.0010
80	1.9970	0.0005

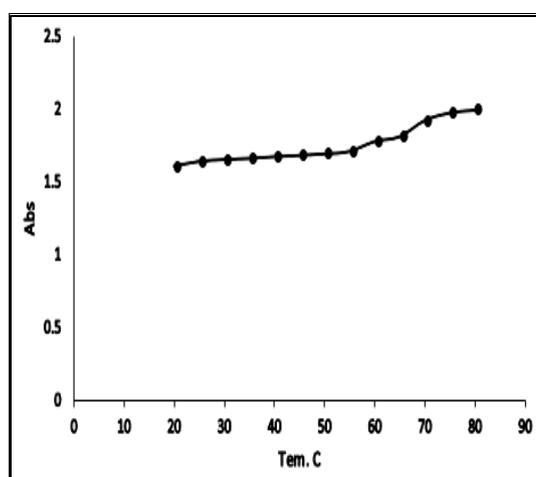


Figure (15): The effect of temp. on the stability of cis-[Pt(4AAP) (NH₃)₂] I₂ complex

Determination of the expiry date of the complexes:

The expiry date of the complexes were determined in 35°C and 20°C for the solutions of the complexes (1x10⁻³M). This study shows that the expiry date at 35°C of the complex cis-[Pt(4AAP) (NH₃)₂] I₂ is 270 hours, while it is for the complex cis-[Pt(4 ASA) (NH₃)₂]I is 420 hours and at 20°C the expiry date is between 410-435 hours for both complexes. Table (7,8,9) and figures (17,18,19,20) show this study.

Table (7): Expiry date data at 35⁰C of the complexes

Tim. hour	Abs. Cis-[Pt(4AAP) (NH₃)₂] I₂	Abs. Cis-[Pt(4 ASA) (NH₃)₂]I
0	2.5970	2.6032
48	2.5600	2.2898
96	2.1970	2.0759
144	2.1430	2.0495
192	1.9660	2.0162
240	1.8760	1.8987
288	1.8320	1.6981
336	1.5390	1.5723
384	1.2530	1.3259
432	1.0100	1.1216

Table (8): Expiry date data at 20⁰C of the complexes

Tim. hour	Abs. Cis-[Pt(4AAP) (NH₃)₂] I₂	Abs. Cis-[Pt(4 ASA) (NH₃)₂]I
0	2.4700	3.1169
48	2.2350	2.5199
96	2.1600	2.3489
144	1.9020	2.1471
192	1.8990	2.0467
240	1.8110	2.0348
288	1.7380	2.0322
336	1.4000	2.0283
384	1.2200	1.8633
432	0.9620	1.2434

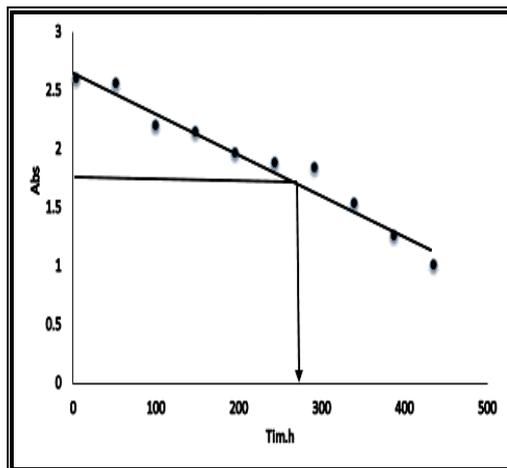


Figure (17): Expiry date for cis-[Pt(4-AAP) (NH₃)₂] I₂ at 35 °C

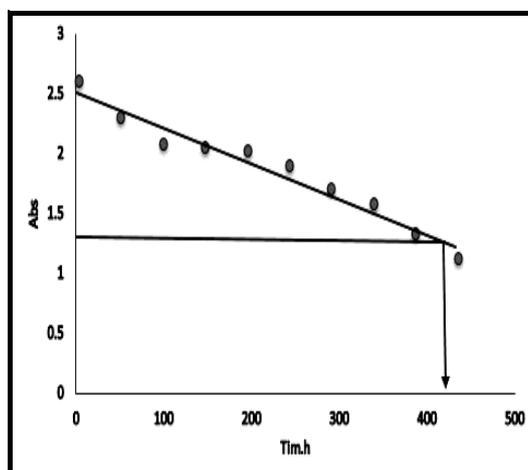


Figure (18):): Expiry date for cis-[Pt(4-ASA) (NH₃)₂]I at 35 °C

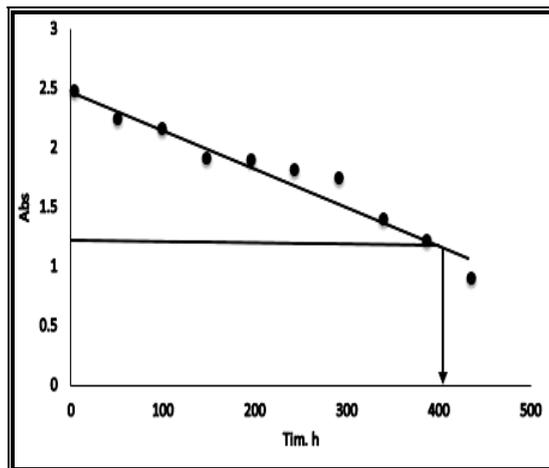


Figure (19): Expiry date for cis-[Pt(4-AAP) (NH₃)₂] I₂ at 20 °C

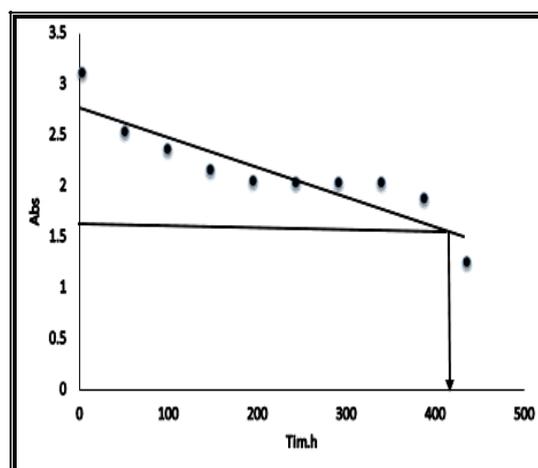


Figure (20): Expiry date for cis-[Pt(4-ASA) (NH₃)₂] I at 20 °C

Table (9): Expiry date of the complexes in hours

Complexes	(t _{1/2}) in 35-40 °C	(t _{1/2}) in 20 °C
Cis-[Pt(4-AAP) (NH ₃) ₂] I ₂	270 h	410 h
Cis-[Pt(4-ASA) (NH ₃) ₂] I	420 h	435 h

CONFLICT OF INTERESTS.

There are non-conflicts of interest.

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الخلاصة

يتضمن هذا البحث تحضير معقد جديد للسز – بلاتين بصيغة الأيودين (Cis – Iodoplatin) وكذلك معقدين جديدين للسز – كاربوبلاتين وهما : cis [Pt(aminosalicylicacid)(NH₃)₂]₂ و cis [Pt (4- aminoantipyrine) (NH₃)₂]₂ واللذان يمكن ان يستخدموا كعقاقير ضد السرطان، وقد حضرا باستخدام الليكاندات العضوية : 4-amino salicylic acid و 4-amino antipyrine . شخّصت هذه المعقدات باستخدام التقنيات الطيفية : UV-Vis و FTIR و XRD و ¹HNMR . وقد درست تأثير التركيز ودرجة الحرارة على التوصيلية الكهربائية لمحاليل هذه المعقدات. وقد حدد تاريخ انتهاء الصلاحية (عمر النصف) لمحاليل هذه المعقدات في ظروف درجة الحرارة 35 و 20 درجة مئوية.